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Solution . To get the requested answer, let us formulate a "stoichiometric" equation (molar quantities) for the reaction: $C_2H_6 + 7O_2 \rightarrow 2CO_2 + 3H_2O$. Each C_2H_6 (ethane) molecule requires 7 oxygen atoms for complete combustion. In molar quantities: 1 mole of $C_2H_6 = 2 \times 12.01 + 6 \times 1.008 = 30.07$ g requires $.7 \times 15.9984$ g = 1.12×10

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An Introduction to Combustion (Burning) Reactions

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